

# Teaching Transparency Worksheet Balancing Chemical Equations Answers

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### Teaching Transparency Worksheet Balancing Chemical

#### **Balancing Equations: Practice Problems - North Allegheny**

Balancing Equations: Answers to Practice Problems 1 Balanced equations (Coefficients equal to one (1) do not need to be shown in your answers)

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A chemical is dissolved in water A chemical is in the liquid state 2, Finish balancing the following equation:  $4 \text{ Al} + 3 \text{ O}_2 \rightarrow 2 \text{ Al}_2\text{O}_3$  + 9 Fe In each of the following formulas, write the total number of atoms present 12 atoms 16 atoms 51 atoms 3 x 10<sup>24</sup> atoms a b 802 23 d 6 X 10 HN<sub>3</sub> 4 Convert the following word equation into a balanced chemical

#### **Worksheet: Writing and Balancing Chemical Reactions**

Worksheet: Writing and Balancing Chemical Reactions 1 Balance the following equations and indicate the type of reaction as formation, decomposition, single

#### **Balancing Word Equations Chapter 9 - My Chemistry Class**

Worksheet Balancing Word Equations Chapter 10 (Remember the following are diatomic: H<sub>2</sub>, N<sub>2</sub>, O<sub>2</sub>, F<sub>2</sub>, Cl<sub>2</sub>, Br<sub>2</sub>, I<sub>2</sub>) The coefficients should add up to the number at the end that is in parenthesis

### TEACHING TRANSPARENCY WORKSHEET The Activity Series

TEACHING TRANSPARENCY WORKSHEET The Activity Series Use with Chapter 9, Section 92.1 For each of the following pairs of elements, underline the one that would replace the other element in a compound a calcium, tin e iron, copper b bromine, fluorine f iodine, chlorine c aluminum, potassium g silver, lead d zinc, sodium 2

### Balancing Chemical Equations - AP Chemistry

Balancing Chemical Equations - Answer Key Balance the equations below: 1)  $1 \text{ N}_2 + 3 \text{ H}_2$

### Balancing Chemical Equations Worksheet 1

Balancing Chemical Equations Worksheet 1 - ANSWERS 1  $2 \text{ H}_2 + \text{O}_2 \rightarrow 2 \text{ H}_2\text{O}$  2  $2 \text{ Na} + \text{Cl}_2 \rightarrow 2 \text{ NaCl}$  3  $\text{N}_2 + \text{O}_4 \rightarrow 2 \text{ NO}_2$  4  $2 \text{ Mg} + \text{O}_2 \rightarrow 2 \text{ MgO}$  5  $2 \text{ H}_2 + \text{O}_2 \rightarrow 2 \text{ H}_2\text{O}$  6  $3 \text{ Ca} + \text{N}_2 \rightarrow \text{Ca}_3\text{N}_2$  7  $2 \text{ Li} + \text{F}_2 \rightarrow 2 \text{ LiF}$  8  $3 \text{ Mg} + \text{N}_2 \rightarrow \text{Mg}_3\text{N}_2$  9  $2 \text{ NH}_3 \rightarrow 3 \text{ N}_2 + 2 \text{ H}_2$  10  $2 \text{ HCl} \rightarrow \text{H}_2 + \text{Cl}_2$  11  $2 \text{ NI}_3 \rightarrow \text{N}_2 + 3 \text{ I}_2$  12  $2 \text{ HI} \rightarrow \text{H}_2 + \text{I}_2$  13

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For each of the following chemical reactions, write a word equation, a skeleton equation, and a balanced chemical equation Be sure to show the state of each reactant and product If you need more help writing formulas or determining the state of a substance, refer to Chapters 7 and 8 and the periodic table on pages 178–179 17

### Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_ TEACHING TRANSPARENCY

22 Chemistry: Matter and Change Teaching Transparency 1 Matter and Change Teaching Transparency Worksheet 2 in what chemical family is it located in the periodic table? \_\_\_\_\_ 11 What is the electron configuration of O<sup>2-</sup> ion? What neutral atom has same configuration, and in what

### Chapter 9: Chemical Reactions

chemical formulas to show the identities and relative amounts of the substances involved in a chemical reaction Balancing Chemical Equations The balanced equation for the reaction between aluminum and bromine, shown in Figure 95, reflects the law of conservation of mass To balance an equation, you must find the correct coefficients for the

### Chemical Reactions - Glencoe

- Write complete ionic and net ionic equations for chemical reactions in aqueous solutions
- Predict whether reactions in aqueous solutions will produce a precipitate, water, or a gas

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### balancing equations worksheet and key 7 23 09

Balancing Equations Worksheet and Key 1 Answer the following questions about the chemical equation shown below:  $2 \text{ H}_2 + \text{O}_2 \rightarrow 2 \text{ H}_2\text{O}$  a) What are the reactants? b) What is the product? c) What do we call the number "2" in front of the H<sub>2</sub> (and H<sub>2</sub>O)? d) Is the reaction balanced? e) Why is there not a coefficient for O<sub>2</sub>?

### Teaching Transparency 19 Answer Key

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### **Redox Reactions - Glencoe**

- Interpret redox reactions in terms of change in oxidation state Lesson Resources Section Focus Transparency 75 and Master Teaching Transparencies 59 -60 and Masters ChemLab and MiniLab Worksheets, p 77 TCR Study Guide for Content Mastery, pp 115-116 TCR Multimedia Resources Chemistry Interactive CD-ROM, Section 201

**Teacher: Ms. Desso Semester: Fall 2011 Class: Physical ...**

Properties of Matter Chemical vs Physical Changes States of Matter Standards 3a, 3b, 3d, Balancing Chemical Equations Prior to teaching the lesson on chemical reactions and energy transfer the students will be shown a demonstration 3:

### **Chemical Reactions**

The chemical equation gives the relative amounts of reactants and products 10 Explain why it is important to reduce coefficients in a balanced equation to the lowest possible whole-number ratio Coefficients in the lowest ratio most clearly indicate the relative amounts of substances in a reaction 11 Analyze When balancing a chemical equation,

### **CHAPTER 19 Oxidation-Reduction Reactions**

Reduction Reactions in which the oxidation state of an element decreases are reduction processes Consider the behavior of chlorine in its reaction with sodium Each chlorine atom accepts an electron and becomes a chloride